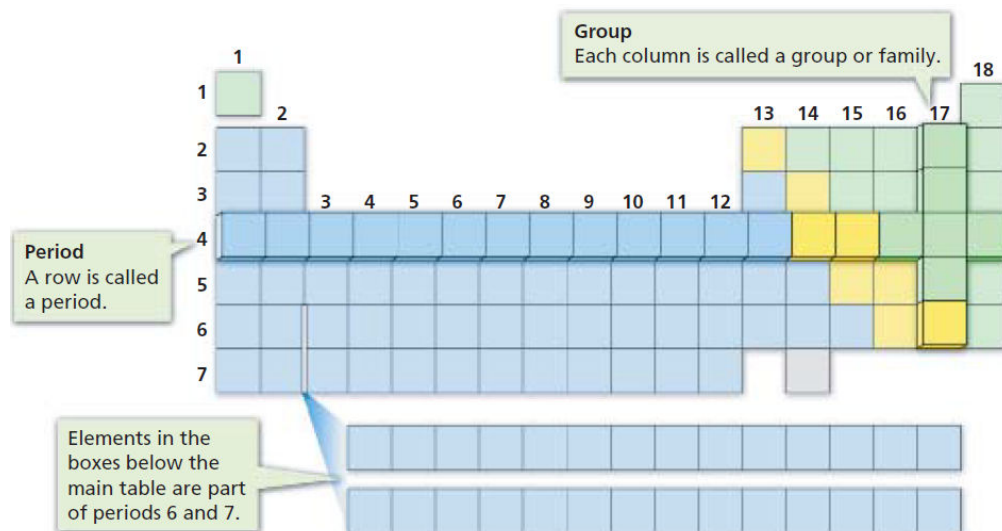


Vocabulary: Atoms & Elements, Part 2



metals:

Metals make up the majority of the elements in the periodic table—more than 75%! They are located on the left side of the periodic table and in the center.

- Group 1 metals are called alkali metals (Hydrogen is the only exception—it is a gas at room temperature and is considered a nonmetal). The alkali metals are the most chemically reactive metals on the periodic table.
- Group 2 metals are called alkaline earth metals.
- The metals in the block of elements in the center of the periodic table (Groups 3-12) are called the transition metals.
- There are other metals in groups 13-15.
- Here are a few important properties of metals:
 - They have a metallic shine, or luster.
 - They are usually solids at room temperature.
 - They are malleable, meaning that they can be hammered, pounded, or pressed into different shapes without breaking.
 - They are ductile, meaning that they can be drawn into thin sheets or wires without breaking.
 - They are good conductors of heat and electricity.

nonmetals:

Nonmetals are on the right side of the periodic table in groups 14-18. There are only 18 elements that fall into this category.

- Here are a few important properties of nonmetals:
 - They rarely have metallic luster.
 - They are usually gases at room temperature.
 - Nonmetallic solids are neither malleable nor ductile.
 - They are poor conductors of heat and electricity.
- The nonmetals found in group 17 are called the halogens (salt making).
- The nonmetals found in group 18 are called the noble gases.

metalloids:

The elements that diagonally stair step between the metals and the nonmetals are called the metalloids (or **semiconductors**). They have properties of both metals and nonmetals.

- The seven elements that fall into this category are: boron (B), silicon (Si), germanium (Ge), arsenic (As), antimony (Sb), tellurium (Te), and polonium (Po). Astatine (At) is sometimes included in this category.