Vocabulary: Temperature of Various Substances

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calorimeter:	An insulated vessel for measuring the amount of heat absorbed or released by a chemical or physical change.
energy:	 Energy is an abstract property something has which enables it to do work. The more energy something has, the more work it can perform. Every kind of energy falls into one of two general categories: kinetic energy or potential energy.
heat:	 The transfer of thermal energy from one system (or body) to another. Heat may be thought of as internal energy in transit. Thermal energy is transferred from the system with the higher temperature to the system with the lower temperature. The greater the heat absorbed by a material, the more rapidly the atoms within the material begin to move, and thus the greater the rise in temperature.
kinetic energy:	 The energy that an object has because of its motion. The kinetic energy is the work done by an external force to bring the body from rest to a particular state of motion.
temperature:	 The measurement of the average kinetic energy of the molecules in an object or system. It can be measured with a thermometer. Provides an indicator of the direction of internal energy flow: when two objects are in contact, internal energy goes from the one at the higher temperature to the one at the lower temperature.
thermal energy:	 The internal energy in a substance due to the movement of the atoms and molecules within that substance. The thermal energy of a substance is equal to the total kinetic energy of its atoms and/or molecules.
thermometer:	An instrument for measuring temperature.
work:	 Work is the energy required to move an object against an opposing force. Work is usually expressed as a force times a displacement. Dropping a stone from a window involves no work, because there is no force opposing the motion (unless you consider air friction) Pushing against a stone wall involves no work, unless the stone wall actually moves.